347.11ml of 0.68 M NaBr solution

534.49 mL of water added to it. Then new volume (347.11+534.49) mL =881.6 mL

According to definition $S_1 V_1 = S_2 V_2$

Where S_1 = 0.68 M V_1 = 347.11 mL V_2 = 881.6 mL

Thus S_2 = 0.267 M

Thus Final concentration = 0.267 M