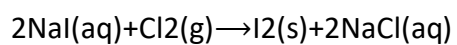


Question #85327, Chemistry / General Chemistry | for completion

Iodine is prepared both in the laboratory and commercially by adding $\text{Cl}_2(\text{g})$ to an aqueous solution containing sodium iodide.



How many grams of sodium iodide, NaI , must be used to produce 27.1g of iodine, I_2 ?

Solution:

$$m(\text{NaI}) = \frac{149.89 \text{ g/mol} \times 2 \times 27.1 \text{ g}}{126.9 \text{ g/mol} \times 2} = 32.00 \text{ g}$$

Answer: 32.00 g

Answer provided by www.AssignmentExpert.com