

Why are the fluorides and oxides of Kr, Xe and Rn the only binary compounds of the noble gases? Why are similar compounds not formed by He and Ne?

The noble gases have full valence electron shells. Valence electrons are the outermost electrons of an atom and are normally the only electrons that participate in chemical bonding. Atoms with full valence electron shells are extremely stable and therefore do not tend to form chemical bonds and have little tendency to gain or lose electrons. However, heavier noble gases such as Rn, Kr, Xe are held less firmly together by electromagnetic force than lighter noble gases such as He, Ne making it easier to remove outer electrons from heavy noble gases.

Sources:

https://en.m.wikipedia.org/wiki/Noble_gas

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