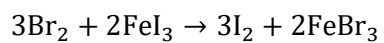


In the reaction  $3\text{Br}_2 + 2\text{FeI}_3 \rightarrow 3\text{I}_2 + 2\text{FeBr}_3$ , how many moles of iodine are produced if 4.7 moles bromine are reacted with excess iron (III) iodide?

**Solution**



$$n(\text{I}_2) = \frac{4.7 \text{ moles} \cdot 3 \text{ moles}}{3 \text{ moles}} = 4.7 \text{ moles}$$

**Answer:**  $n(\text{I}_2) = 4.7 \text{ moles}$

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