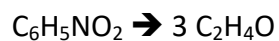


Question #84765, Chemistry / General Chemistry

Assume you have 2.75 moles of $C_6H_5NO_2$. How many grams of C_2H_4O could you make?

Solution



According to the general scheme of the reaction 1 mole of $C_6H_5NO_2$ gives 3 moles of C_2H_4O .

So 2.75 moles of $C_6H_5NO_2$ give 8.25 moles of C_2H_4O .

Find the mass of C_2H_4O :

$$m = M \times v = 44.05 \times 8.25 = \mathbf{363.4 \text{ (g)}}.$$

Answer

363.4 g of C_2H_4O could be made from 2.75 moles of $C_6H_5NO_2$.

Answer provided by www.AssignmentExpert.com