

What is the buffer concentration and PH of a mixture of 0.042M NaH₂PO₄ (pK=6.86) and 0.058 M Na₂HPO₄?

Solution.

$$\text{pH} = \text{pK} + \lg C_b/C_{ac} = 6.86 + \lg(0.058/0.042) = 7$$

$$[\text{H}^+] = 10^{-7} \text{ mole/l}$$

Answer. pH = 7; [H⁺] = 10⁻⁷

Answer provided by www.AssignmentExpert.com