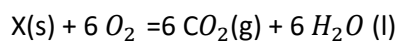


Answer on Question 83873 in General Chemistry.



$$\Delta H = -3840 \text{ kJ/mol}$$

$$\Delta H(CO_2) = -393.5 \text{ kJ/mol}$$

$$\Delta H(H_2O) = -285.8 \text{ kJ/mol}$$

$$\Delta H(X) = ?$$

$$\Delta H = 6 \times \Delta H(CO_2) + 6 \times \Delta H(H_2O) - \Delta H(X)$$

From which

$$\Delta H(X) = \Delta H - 6 \times \Delta H(CO_2) - 6 \times \Delta H(H_2O) = -3840 - 6 \times (-393.5) - 6 \times (-285.8) = 235.8 \text{ kJ/mol}$$

Answer provided by www.AssignmentExpert.com