

## Answer on Question #83763 - Chemistry - General Chemistry

Question:

4. Consider the inter-nuclear distance of  $\text{CuCl}_2$  crystal is 1.31 angstrom. Calculate the ionic radius of  $\text{Cu}^{2+}$  and  $\text{Cl}^-$  separately

**Solution:**

$$r = d/2 = 1.31/2 = 0,655 \text{ angstrom} - \text{Cu}^{2+};$$

$$r(\text{Cl}^-) = 0.655 / 2 = 0.3275.$$

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