

how many moles of H are there in 120.5mg of ammonium carbonate

Solution.

$$n((\text{NH}_4)_2\text{CO}_3) = 1.255 \text{ mmoles}$$

$$n(\text{H}) = 1.255 \times 8 = 10.04 \text{ mmoles} = 10.04 \times 10^{-3} \text{ moles}$$

Answer. 10.04×10^{-3} moles (≈ 0.01 mole)

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