Calculate the number of mL of 2.00 M HNO solution required to react with 216 of Ag.

## Solution.

$\mathrm{n}(\mathrm{Ag})=216 / 108=2$ moles
$\mathrm{Ag}+2 \mathrm{HNO} 3=\mathrm{AgNO} 3+\mathrm{NO} 2+\mathrm{H} 2 \mathrm{O}$
$\mathrm{N}(\mathrm{HNO} 3)=4$ moles
$\mathrm{V}(\mathrm{HNO})=4 / 2=2 \mathrm{~L}=2000 \mathrm{~mL}$
Answer. 2000 mL.

Answer provided by www.AssignmentExpert.com

