Calculate the number of mL of 2.00 M HNO3 solution required to react with 216 of Ag.

## Solution.

$$n(Ag) = 216/108 = 2 \text{ moles}$$

$$Ag + 2HNO3 = AgNO3 + NO2 + H2O$$

$$N (HNO3) = 4 moles$$

$$V (HNO3) = 4/2 = 2 L = 2000 mL$$

Answer. 2000 mL.

Answer provided by www.AssignmentExpert.com