

$$PV = nRT$$

$$R = 0.08206 \frac{L \cdot atm}{mol \cdot K}$$

$$T_{(K)} = T_{(^{\circ}C)} + 273.15 = 208.35 \text{ K}$$

$$V = \frac{nRT}{P} = \frac{0.0605 \text{ mol} \cdot 0.08206 \frac{L \cdot atm}{mol \cdot K} \cdot 208.35 \text{ K}}{0.3826 \text{ atm}} = 2.7 \text{ L}$$

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