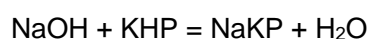


Question #83532, Chemistry / General Chemistry | for completion

A NaOH solution has a concentration of about 0.14 M. How many grams of KHP would have to be used to require approximately 25 mL of the NaOH solution to reach the endpoint? Show your work.

Answer:



$$C_M = n/V \text{ and } n = C_M \times V$$

$$n = 0.14 \times 0.025 = 0.0035 \text{ mol}(\text{NaOH})$$

$$n = m/M_r, m = n \times M_r$$

$$m = 0.0035 \times 204 = 0.714 \text{ g (KHP)}$$

Answer provided by www.AssignmentExpert.com