

Determine the number of moles of oxygen atoms in: 5.02 mol H₂O₂.

Solution:

Each molecule of hydrogen peroxide contains 2 atoms of oxygen, so there will be double amount of oxygen atoms per each amount of hydrogen peroxide. In other words:

$$n(\text{O}) = 2 * n(\text{H}_2\text{O}_2) = 5.02 \text{ mol} * 2 = 10.04 \text{ mol}$$

(n – number of moles)

Answer:

There are 10.04 moles of oxygen atoms in 5.02 moles of H₂O₂.

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