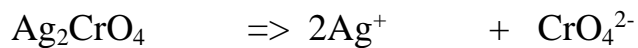


A. Calculate K_{sp} of Ag_2CrO_4 . Organize your answer so that your solution is logical. Make sure that your K_{sp} has correct significant figures.

B. From the experiment $Cu = 0.00016$ mol. Also $Ag_2CrO_4 = 0.0016$ mol

b) Calculate percentage error.



$$K_{sp} = [Ag^+]^2 [CrO_4^{2-}] = (0.00032 \text{ mol})^2 \cdot 0.0016 \text{ mol} = 1.6 \cdot 10^{-11}$$

$$K_{sp} (Ag_2CrO_4) = 1.1 \cdot 10^{-12} \text{ at } 298 \text{ K}$$

$$\% \text{ error} = ((\text{experimental} - \text{theoretical}) / \text{theoretical}) \cdot 100$$

$$\% \text{ error} = (1.6 \cdot 10^{-11} - 1.1 \cdot 10^{-12}) / 1.1 \cdot 10^{-12} \cdot 100 = 14\%$$

Answer provided by www.AssignmentExpert.com