

## Question #82758, Chemistry / General Chemistry | for completion

A monatomic ion with a charge of -2 has an electronic configuration of  $1s^2 2s^2 2p^6 3s^2 3p^6$ .

This ion is a(n) .

What is the chemical symbol of the noble gas this ion is isoelectronic with? .

What is the formula of the ion? .

Answer:

$X^{2-} 1s^2 2s^2 2p^6 3s^2 3p^6$  it is Sulfide ion, **because atomic number is 16(18 - 2).**

**Noble gas this ion is isoelectronic with Ar.**

**Formula of the ion  $S^{2-}$**

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