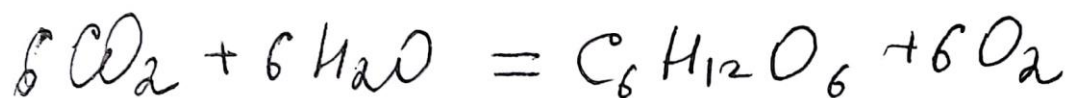


If a plant produces 5.81 mol of $C_6H_{12}O_6$, how many moles of CO_2 are needed?

Solution:



$$n(C_6H_{12}O_6) = 5.81 \text{ moles.}$$

$$\frac{n(CO_2)}{6} = \frac{5.81 \text{ moles}}{1}$$

$$n(CO_2) = 34.86 \text{ moles}$$

$$\text{Answer : } n(CO_2) = 34.86 \text{ moles.}$$