If a plant produces 5.81 mol of C6H12O6, how many moles of CO2 are needed? Solution:

$$SCO_2 + 6H_{2O} = C_6H_{12}O_6 + 6O_2$$
 $R(C_6H_{12}O_6) = 5.81$  moles.

 $R(CO_2) = \frac{5.81}{6}$  moles

 $R(CO_2) = \frac{34.86}{1}$  moles

Answer:  $R(CO_2) = 34.86$  moles.

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