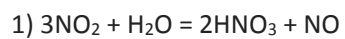


Suppose that 4.5 mol NO₂ and 0.70 mol H₂O combine and react completely. Which reactant is in excess?

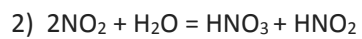
Solution.

Depending on conditions



$$n(\text{NO}_2) = 0.7 \cdot 3 = 2.1 \text{ moles}$$

$$n(\text{H}_2\text{O}) = 0.70 \text{ moles}$$



$$n(\text{NO}_2) = 0.7 \cdot 2 = 1.4 \text{ moles}$$

$$n(\text{H}_2\text{O}) = 0.70 \text{ moles}$$

Answer: NO₂

Answer provided by www.AssignmentExpert.com