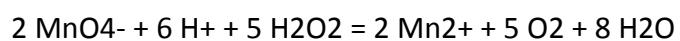
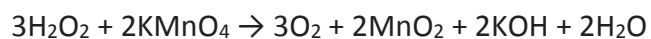


Based on the stoichiometry of the reaction, what is the ratio of  $\text{MnO}_4^-$  to  $\text{H}_2\text{O}_2$  consumed in the reaction?

- a)  $5\text{MnO}_4^- : 2\text{H}_2\text{O}_2$
- b)  $1\text{MnO}_4^- : 1\text{H}_2\text{O}_2$
- c)  $2\text{MnO}_4^- : 10\text{H}_2\text{O}_2$
- d)  $2\text{MnO}_4^- : 5\text{H}_2\text{O}_2$

**Solution.**



**Answer. d)**

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