

Based on the number of moles of old H_2O_2 consumed in the reaction, how many grams of H_2O_2 were in the hydrogen peroxide solution? The molar mass of H_2O_2 is 34.0147 g/mol. Choose the closest answer.

a) 0.306g

b) 0.215g

c) 0.651g

d) 0.895g

If we take 1L of old H_2O_2

$$C(\text{H}_2\text{O}_2) = 34,0147 \text{ g/mol}$$

$$n(\text{H}_2\text{O}_2) = 34,0147 \text{ mol}$$

$$M(\text{H}_2\text{O}_2) = 34 \text{ g/mol}$$

$$m(\text{H}_2\text{O}_2) = 34,0147 \text{ mol} / 34 \text{ g/mol} = 1,0004 \text{ g}$$

The closest answer is d)0.895g.

Answer provided by www.AssignmentExpert.com