Enter the balanced ionic and net ionic equation of:
H2SO3(aq)+2NH3(aq) --> (NH4)2SO3 (aq)
Solution:
$\begin{aligned} \mathrm{H}_{2} \mathrm{RO}_{3}(\mathrm{aq})+2 \mathrm{NH}_{3}(\mathrm{aq}) & \rightarrow\left(\mathrm{M}_{4}\right)_{2} \mathrm{NO}_{3}(\mathrm{aq}) \\ \mathrm{H}_{2} \mathrm{OO}_{3}(\mathrm{aq})+2 \mathrm{H}_{3}(\mathrm{aq}) & \rightarrow 2 \mathrm{NH}_{5}^{+}+\mathrm{PO}_{3}^{2-}(\mathrm{aq})\end{aligned}$
Strong electrolytes, dissociated virtually completely, are recorded in the forme of ions Low-dissociated (weak electrolytes), sparingly
soluble and gaseous substances are recorded in the form of molecules.

