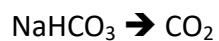


Answer on Question #81431, Chemistry / General Chemistry

A teaspoonful of baking soda contains a mass of 11g of sodium hydrocarbonate. Calculate the mass of carbon dioxide that could be made from 11g of sodium hydrocarbonate.

Solution



According to the scheme 1 mole of sodium hydrocarbonate produces 1 mole of CO_2 .

Find the amount of sodium hydrocarbonate:

$$v = \frac{11}{84} = 0.131 \text{ (mole)}$$

Find the mass of 0.131 mole of CO_2 :

$$m = 0.131 \times 44 = \mathbf{5.764 \text{ (g)}}$$

Answer

5.764 g of carbon dioxide could be made from 11g of sodium hydrocarbonate.

Answer provided by www.AssignmentExpert.com