Answer on the question 81385 – Chemistry – General Chemistry

$P(partial) = nRT/V$, where n – moles, $R = 0.0831$ l*atm/moles*K, T – temperature, V – volume;

$V(CO_2) = n*V(const) = 8 * 22.4 = 179.2$ l; $V(O_2) = n*V(const) = 2*22.4 = 44.8$l

$P(partial) = 0.0831*273*352/(179.2+44.8) = 0.8$ atm.

Answer: 0.8 atm.