The mass ratio of sodium to fluorine in sodium fluoride is 1.21:1. A sample of sodium fluoride produced 29.3 g of sodium upon decomposition. How much fluorine (in grams) was formed?

Solution:

Let **x** be the mass of formed fluorine in grams. Then, according to mass ratio, mass of sodium will be 1.21***x.** Also, sodium mass from the task is 29.3 g. Thus, final equation will look like:

1.21x=29.3;

x= 29.3/1.21;

x=24.2 g

Answer:

The mass of formed fluorine is 24.2 g.