

## Question #80760, Chemistry / General Chemistry

What is the pH of solution containing  $2.3 \times 10^{-2}$  mol/l of  $H^+$  ions?

**Solution:**  $pH = -\log_{10}[H^+]$

$$pH = -\log_{10}[2.3 * 10^{-2}] = 1.6$$

**Answer:** 1.6

**Source:** <http://www.usetute.com.au/phhcalcs.html>

Answer provided by AssignmentExpert.com