Answer on Question 80615 – Chemistry – General Chemistry

In true variant, formules are:

1)
$$H_2O_2 + I^- = H_2O + IO^-$$
 (because there is not H_2OH)

2)
$$H_2O_2 + IO^- = H_2O + O_2 + I^-$$

The reaction $2H_2O_2 = H_2O + O_2$ make in 2 steps. And the catalyst of this is I⁻ (we have it at the begining of reaction and at the end), which we haven't in the variants (IO⁻ is an intermediate).

So, the answer is E.

Amswer: E).

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