

Which of the following compounds has the greatest percentage of oxygen?

Given the atomic masses:

C = 12, O = 16, N = 14, S = 32, and Mg = 24

Solution:

$$\% (O) = \frac{n \times 16}{M(\text{compound})}$$

1) First of all, calculate the M of all the molecules

2) After that, use the formula

$$\% (O) = \frac{n \times 16}{M(\text{compound})}, \text{ where}$$

$n \times 16$ – the total atomic mass of O in each molecule.

Answer provided by www.AssignmentExpert.com