

Task #80570

Uranium(Gallium?) has two isotopes, Gallium-69 and Gallium-71. How many neutrons are present in an atom of Gallium 71?

Solution.

Probably, a typo in the task. I will solve this task with the condition that the task is Gallium, not Uranium. The number of neutrons in an isotope or any element is:

$n^{\circ} = A_r - Z$, where A_r – atomic mass, Z - charge of an atom.

Answer:

40 n°

The charge of the Gallium isotope is 31, and the atomic-molecular mass is 71, which means that the number of neutrons in the Gallium-71 isotope is 40.

$n^{\circ} = 71 - 31 = 40 n^{\circ}$

Answer provided by www.AssignmentExpert.com