$n = m/M = N/N_A$

N – how many atoms or molecules are in a given substance.

N_A – Avogadro's number

n – how many moles of a given substance

$n = m/M = 4.77/18 = 0.265$ moles

$M(H_2O) = 18$

$N = n * N_A = 0.265$ moles $* 6.022 \times 10^{23} = 1.596 \times 10^{23}$ molecules