

## Answer on Question #80139 - Chemistry - Physical Chemistry

Question:

A solution of 0.01M concentration of  $\text{NH}_4\text{OH}$  is 2.6% dissociated, calculate the pH of the solution.

**Solution:**

$$0.01 \times 0.026 = 0.00026 \text{ M};$$

$$\text{pOH} = -\lg[\text{OH}] = 3.58;$$

$$\text{pH} = 14 - \text{pOH} = 14 - 3.58 = 10.41.$$

**Answer:** 10.41.

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