

#80060 Chemistry, Other

At 25°C and 1 atm pressure, ethylene glycol has a density of 1.11 g/mL. Calculate the volume of 0.025 kg of ethylene glycol.

Answer:

$$\rho = \frac{m}{V}$$

$$V = \frac{m}{\rho}$$

$$V(C_2H_6O_2) = \frac{0.025 \times 1000}{1.11} = 22.52 \text{ ml} = 0.02 \text{ L}$$

Answer provided by AssignmentExpert.com