V = 8.50 L = 8 500 ml

m (C_8H_{18}) = d *V = 0.758 g/ml * 8 500 ml = 6 443 g

6 443 g n moles x g z g

2 C_8H_{18} + 25 O_2 = 16 CO_2 + 18 H_2O

2*114 g 25 moles 16*44 18*18

M (C_8H_{18}) = 114, M (CO_2) = 44, M (H_2O) = 18.

n = (6 443 *25)/ (2 *114) = 706.47 moles – amount of O_2 needed for the complete combustion of gasoline in one day.

z = (6 443 *18*18) / (2*114) = 9155.84 g of H_2O formed per day

x = (6 443 *16*44) / (2*114) = 19894.2 g of CO_2 formed per day

19 894.2 *2 = 39788.4 g of CO_2 formed in two days.