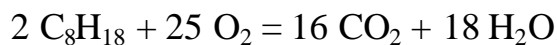


$$V = 8.50 \text{ L} = 8\,500 \text{ ml}$$

$$m(\text{C}_8\text{H}_{18}) = d * V = 0.758 \text{ g/ml} * 8\,500 \text{ ml} = 6\,443 \text{ g}$$

$$6\,443 \text{ g} \quad n \text{ moles} \quad x \text{ g} \quad z \text{ g}$$



$$2 * 114 \text{ g} \quad 25 \text{ moles} \quad 16 * 44 \quad 18 * 18$$

$$M(\text{C}_8\text{H}_{18}) = 114, \quad M(\text{CO}_2) = 44, \quad M(\text{H}_2\text{O}) = 18.$$

$n = (6\,443 * 25) / (2 * 114) = 706.47$ moles – amount of O_2 needed for the complete combustion of gasoline in one day.

$$z = (6\,443 * 18 * 18) / (2 * 114) = 9155.84 \text{ g of H}_2\text{O formed per day}$$

$$x = (6\,443 * 16 * 44) / (2 * 114) = 19894.2 \text{ g of CO}_2 \text{ formed per day}$$

$$19\,894.2 * 2 = 39788.4 \text{ g of CO}_2 \text{ formed in two days.}$$