- 1. Consider the reaction when aqueous solutions of magnesium chloride and potassium phosphate are combined. The net ionic equation for this reaction is:
- 2. Write a net ionic equation for the reaction that occurs when aqueous solutions of barium hydroxide and hypochlorous acid are combined.

Solution:

1. $2 K_3PO_4(aq) + 3 MgCl_2(aq) = Mg_3(PO_4)_2(s) + 6 KCl(aq);$

Type of reaction: ion exchange.

$$6K^{+} + 2PO_{4}^{3-} + 3Mg^{2+} + 6CI^{-} = Mg_{3}(PO_{4})_{2}(s) + 6K^{+} + 6CI^{-};$$

 $2PO_{4}^{3-} + 3Mg^{2+} = Mg_{3}(PO_{4})_{2}(s).$

2.
$$HCIO(aq) + Ba(OH)_2(aq) = Ba(OH)OCI(s) + H_2O(aq);$$

 $HCIO + Ba^{2+} + 2OH^{-} = Ba(OH)^{+} + OCI^{-} + H_2O.$

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