

Answer on Question #78563 - Chemistry - Physical Chemistry

Question:

14. The combustion enthalpy of Cellobiose $C_{12}H_{22}O_{11}$ is -5401.5 kJ/mole. Find the enthalpy of formation.

Solution:

$$\Delta H_c = \Delta H_f(\text{reagents}) - \Delta H_f(\text{products}) = -5401.5 \text{ kJ/mole};$$

$$\Delta H_f = \Delta H_f(\text{products}) - \Delta H_f(\text{reagents}) = -\Delta H_c = 5401.5 \text{ kJ/mole}.$$

Answer: 5401.5 kJ/mole.