According to Raoult's law n equation form, for a mixture of liquids A and B, this reads:

$$p_A = x_A \times P_A^o$$
$$p_B = x_B \times P_B^o$$

In this equation, P_A and P_B are the partial vapour pressures of the components A and B.

And

The total vapour pressure of the mixture is equal to the sum of the individual partial pressures.

$$\text{Total vapour pressure} = p_A + p_B = P$$

Also x_A and x_B are the mole fractions of A and B. That is exactly what it says it is - the fraction of the total number of moles present which is A or B.

You calculate mole fraction using, for example:

$$x_A = \frac{\text{moles of A}}{\text{total number of moles}} =$$

From this

$x_A+x_B=1$

$P=P_A*x_A+P_B*x_B=P_A*x_A+P_B-P_B*x_A$

$x_A=(P-P_B)/P_A-P_{B=(}286-395)/(96-395)=0.3645=n_A/(n_A+n_B)$; n-number of moles

$n_{A=}(n_A+n_B)*0.3645$

From the condition, we know that mixture we have equal masses C6H6(we call it a component B) and compound X(we call it a component A). It follows that $M_A*n_A=M_B*n_B=>n_{A=}M_B*n_B/M_A$. Combining these two equations, we obtain $M_B*n_B/M_A=(M_B*n_B+M_A*n_B)/M_A*0.3645$
$=>M_B*n_B=M_B*n_B*0.3645+M_A*n_B*0.3645$
$=>M_{B=}M_B*0.3645+M_A*0.3645$

$=>M_A=(M_B-M_B*0.3645)/0.3645=((6*12+6)-(6*12+6)*0.3645)/0.3645=136$