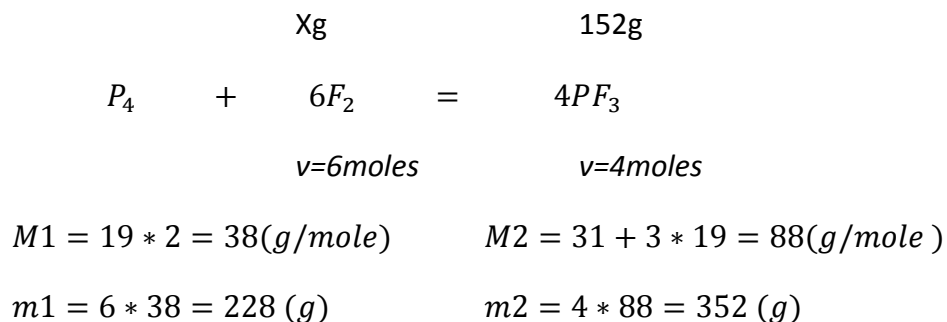


Question #78326, Chemistry / General Chemis try

In the reaction between red phosphorus (P₄) and fluorine, phosphorus trifluoride is produced. If the percent yield is 82.6%, what mass of fluorine is needed to produce 152 g of phosphorus trifluoride?

Solution



$$\frac{Xg}{228g} = \frac{152g}{352g};$$

$$x = \frac{228 * 152}{352} = 98.45g \text{ (100\%yield);}$$

$$m = 98.45 * \frac{100}{82.6} = \mathbf{119.2(g)};$$

Answer provided by AssignmentExpert.com