

Question #78241, Chemistry / General Chemistry

In the equation  $\Delta H_o = \sum \Delta H_{fo} (\text{products}) - \sum \Delta H_{fo} (\text{reactants})$ ,  $\Delta H_o$  stands for \_\_\_\_\_.

- A. standard change of enthalpy
- B. standard change of entropy
- C. change in heat
- D. change in hydrogen bonds

**Solution:** It is the equation for the standard enthalpy change of formation (originating from Enthalpy's being a State Function).

**Answer:** A

**Source:** <http://www.grandinetti.org/enthalpy>

Answer provided by AssignmentExpert.com