Solution

0.318g   60ml

$v = m/M = 0.318/106 = 0.003$moles   x moles

Na_2CO_3 + $2HCl = 2NaCl + H_2O + CO_2$

v=1mole   v=2moles

$M = 2x23 + 12 + 3x16 = 106(g/mole)$ $M = 1 + 35,5 = 36,5(g/mole)$

Proportionally

$$\frac{0.003moles}{1} = \frac{x}{2moles};$$

$$x = \frac{0.003x2}{1} = 0,006moles;$$

Molarity of the acid:

$$\frac{0.006moles}{60ml} = \frac{xmoles}{1000ml};$$

$$x = \frac{0.006x1000}{60} = 0,1M.$$