

**Question #77045, Chemistry / Physical Chemistry**

Which of the following processes would be expected to have a positive  $\Delta S$  value?

- A)  $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) = 2\text{NO}_2(\text{g})$
- B)  $\text{Cl}_2(\text{g}) + \text{Br}_2(\text{g}) = 2\text{ClBr}(\text{g})$
- C)  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) = 2\text{NH}_3(\text{g})$
- D)  $\text{H}_2\text{O}(\text{l}) = \text{H}_2\text{O}(\text{s})$
- E)  $\text{CaCO}_3(\text{s}) = \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$

**Answer:**



Increase of disorder, gas as product, only solid compound as reagent.

Answer provided by AssignmentExpert.com