

Answer on Question #75861, Chemistry / General Chemistry

Calculate the mass of the solute in 750.0 mL of CaCl_2 solution which is 0.500 M

Solution

Find the amount of CaCl_2 in the solution:

$$n = C \times V = 0.5 \times 0.75 = 0.375 \text{ (mol)}$$

Find the mass of CaCl_2 ($M = 111 \text{ g/mole}$)

$$m = 0.375 \times 111 = \mathbf{41.63 \text{ (g)}}$$

Answer

The mass of CaCl_2 in the solution is **41.63 g**.

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