

Question #75128, Chemistry / Other / Completed

Find 2.28×10^5 kg NH₃ in liters at STP

Solution:

$$n = 2.28 \times 10^5 \text{ kg} / 17.031 \text{ kg/kmol} = 13387.35 \text{ kmol}$$

$$V = n \cdot V_m = 13387.35 \text{ kmol} \times 22.4 \text{ m}^3/\text{kmol} = 299876.7 \text{ m}^3 = 299876700 \text{ L} \approx 3 \times 10^8$$

Answer: $3 \cdot 10^8$ L.

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