

Answer on Question 75124 in General Chemistry

$$c_M(\text{CaCl}_2) = 0.30 \text{ M}$$

$$V = 2.0 \text{ L}$$

$$n(\text{CaCl}_2) = ?$$

Solution: To find the molarity we use formula

$$c_M = \frac{n(\text{CaCl}_2)}{V} \text{ where do we find } n(\text{CaCl}_2) = V \times c_M = 2.0 \times 0.30 = 0.60 \text{ mol}$$

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