

Question #74917, Chemistry / Other / Completed

Rutherford's experiments included deflection of D-particles. As a result of its experiments, what did he conclude about the atom?

Answer:

Rutherford carried out the experiment, which demonstrated the nuclear nature of atoms by deflecting alpha particles passing through a thin gold foil. It was Rutherford's interpretation of this data that led him to formulate the Rutherford model of the atom in 1911 – that a very small charged nucleus, containing much of the atom's mass, was orbited by low-mass electrons.

