

Answer on Question 74011- General Chemistry

Given wavelength of photon 3.6 angstrom

$$\lambda = 3.6 \text{ \AA} = 3.6 \times 10^{-10} \text{ m}$$

$$V(\text{photon}) = V(\text{light}) = 3 \times 10^8 \text{ m/s}$$

$$.m = \frac{h}{\lambda V} = \frac{6.626 \times 10^{-34} \text{ kg m}^2 \text{ s}^{-2}}{3.6 \times 10^{-10} \text{ m} \times 3 \times 10^8 \text{ m/s}} = 6.135 \times 10^{-29} \text{ kg}$$