

Answer on Question # 73493 - Chemistry - Physical Chemistry

Explain the position of 1A and 2B elements on the periodic table?

Answer:

Position of an element in periodic table depends upon its electronic configuration and properties. All of the elements in group 1A have one valence electron, their valence shell electronic configuration is s^1 . Alkali metals belong to this group of the periodic table, which includes lithium (Li), sodium (Na), potassium (K), rubidium (Rb), cesium (Cs), and francium (Fr). The elements of group 1A have an oxidation state of +1.

Group 2B of the Periodic table of elements is represented by zinc (Zn), cadmium (Cd), mercury (Hg), and copernicium (Cn). Their valence shell electronic configuration is $d^{10}s^2$. The elements of group 2B have an oxidation state of +2, in which the ions have the rather stable d^{10} electronic configuration, with a full sub-shell.

Answer provided by AssignmentExpert.com