

Answer on Question #72468 - Chemistry - General Chemistry

Question:

can you help identify the limiting agent for a reaction equation and calculate the number of moles of product formed, and the number of moles of excess reagent remaining after the equation it is



Solution:

The limiting agent for a reaction equation is A.

Of 2 mole of substance A and 3 mole of substance C₁₂, 2 mole of products (substances of A₁C₁₃) are formed.

Since the amount of substance C₁₂ with the reaction equation is greater than the amount of substance A, then Substance C₁₂ will be in excess, and after passing the reaction, a certain amount of this substance will remain. Determine how much left matter C₁₂: Substance A was 2 mole, then 2 mole of C₁₂ would be required for the reaction to proceed. Consequently, after the reaction, $3 - 2 = 1$ mole of C₁₂.

Answer provided by AssignmentExpert.com