

#72435 Chemistry, Other

A 0.10 M solution of HF has a pH of 2.10. Calculate the  $K_a$  of HF.

**Answer:**

$$-\log K_a = -\log [H_3O^+]$$

$$pK_a = pH$$

$$K_a = 10^{-pK_a}$$

$$K_a = 10^{-2.10}$$

$$K_a = 0.0076$$

Answer provided by AssignmentExpert.com