#72435 Chemistry, Other

A 0.10 M solution of HF has a pH of 2.10. Calculate the Ka of HF.

Answer:

$-\log K_a = -\log [H_3O^+]$

$pK_a = pH$

$K_a = 10^{-pKa}$

$K_a = 10^{-2.10}$

$K_a = 0.0076$