pH = 3,18

C_{HAn} = 1,00 M = 1,00 mol/L

pH = - [lgH$^+$]

[H$^+$] =- 10 $^{-3,18}$ =6,6*10^{-4} mol/L

suppose that we are dealing with monobasic acid.

HAn↔ H$^+$ + An$^-$
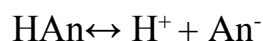
$K_a = \frac{[H+][A-]}{[HA]} = \frac{[H+][H+]}{CHAn-[H+]} \approx \frac{[H+][H+]}{CHAn}$ = (6,6*10^{-4} mol/L)^2 /1,00 mol/L =

= 4,36* 10^{-7} mol/L.