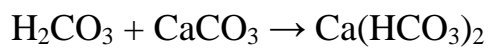


$$K_{(\text{H}_2\text{CO}_3)} = 4,45 \times 10^{-7}$$



$$\text{pH} = -\lg K_{(\text{H}_2\text{CO}_3)} + \lg \frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]}$$

$$\lg \frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = \text{pH} + \lg K_{(\text{H}_2\text{CO}_3)}$$

$$\lg \frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 8.50 - 6.35 = 2.15$$

$$\frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 141.25$$

$$\lg \frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 8.25 - 6.35 = 1.9$$

$$\frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 79.43$$

$$\lg \frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 8.00 - 6.35 = 1.65$$

$$\frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 44.67$$

$$\lg \frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 7.75 - 6.35 = 1.4$$

$$\frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 25.12$$

$$\lg \frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 7.00 - 6.35 = 0.65$$

$$\frac{[\text{CaCO}_3]}{[\text{H}_2\text{CO}_3]} = 4.47$$

Answer provided by AssignmentExpert.com