

## Answer on Question #72058, Chemistry / General Chemistry

How many moles of CO<sub>2</sub> are contained in 6.51 L at 35 °C and 603 torr?

### Solution

PV = nRT, where P – pressure, V – volume, R – gas constant, T – temperature, n – amount.

$$n = \frac{PV}{RT}$$

35°C = 308.15 K, R = 62.363 L×torr/K×mole.

$$n = \frac{603 \times 6.51}{308.15 \times 62.363} = \mathbf{0.204 \text{ mole}}$$

### Answer

**0.204** moles of CO<sub>2</sub> are contained in 6.51 L at 35 °C and 603 torr.

Answer provided by AssignmentExpert.com