Answer on Question #71974, Chemistry / General Chemistry :

How many moles of ammonium hydroxide do you have in a 58.0 gram sample?

Solution.

$m(NH_4OH) = 58.0g$

$M(NH_4OH) = 35.05g/mol$

$v(NH_4OH) - ?$

We have in a 58.0 gram of ammonium hydroxide moles:

$$v = \frac{m(NH_4OH)}{M(NH_4OH)} = \frac{58.0g}{35.05g/mol}$$

$v(NH_4OH) = 1.655mol$

Answer: $v(NH_4OH) = 1.655mol$ **.**